

## Chemistry Section 1 Review Stoichiometry Answers

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**Chemistry Section 1 Review Stoichiometry**  
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This means that moles of different compounds contain different weights. For example, 2 moles of Na = 2 x 22.989 g = 45.98g while 1 mole of Cl = 1 x 35.453 g = 35.453 g Cl. This makes the sodium react completely with chlorine. 2g of sodium would react with = (35.453/45.978) x 2 = 1.542 g Cl

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The stoichiometry of a balanced chemical equation identifies the maximum amount of product that can be obtained. The stoichiometry of a reaction describes the relative amounts of reactants and products in a balanced chemical equation. A stoichiometric quantity of a reactant is the amount necessary to react completely with the other reactant (s).

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A reaction stoichiometry problem in which you are given the mass of B and asked to calculate the number of moles of C produced, you are solving a mass-mole problem The reactant that controls the amount of product formed in a chemical reaction is called the

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Modern Chemistry 77 Stoichiometry CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 \_\_\_\_ The actual yield of a reaction is 22 g and the theoretical yield is 25 g

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